

## ***Common Ions with Multiple Oxidation States***

<b>Cations: Variable Valence</b>		
<b>Ion</b>	<b>Systematic Name</b>	<b>Alternative Name</b>
$\text{Cu}^+$	Copper (I)	Cuprous
$\text{Cu}^{2+}$	Copper (II)	Cupric
$\text{Hg}^+$	Mercury (I)	Mercurous
$\text{Hg}^{2+}$	Mercury (II)	Mercuric
Note: Mercury (I) ions always occur bound together to form $\text{Hg}_2^{2+}$ ions.		
$\text{Au}^+$	Gold (I)	Aurous
$\text{Au}^{3+}$	Gold (III)	Auric
$\text{Co}^{2+}$	Cobalt (II)	Cobaltous
$\text{Co}^{3+}$	Cobalt (III)	Cobaltic
$\text{Cr}^{2+}$	Chromium (II)	Chromous
$\text{Cr}^{3+}$	Chromium (III)	Chromic
$\text{Fe}^{2+}$	Iron (II)	Ferrous
$\text{Fe}^{3+}$	Iron (III)	Ferric
$\text{Mn}^{2+}$	Manganese (II)	Manganous
$\text{Mn}^{4+}$	Manganese (IV)	Manganic
$\text{Pt}^{2+}$	Platinum (II)	Platinous
$\text{Pt}^{4+}$	Platinum (IV)	Platinic
$\text{Sn}^{2+}$	Tin (II)	Stannous
$\text{Sn}^{4+}$	Tin (IV)	Stannic
$\text{As}^{3+}$	Arsenic (III)	Arsenous
$\text{As}^{5+}$	Arsenic (V)	Arsenic
$\text{Sb}^{3+}$	Antimony (III)	Antimonous
$\text{Sb}^{5+}$	Antimony (V)	Antimonic