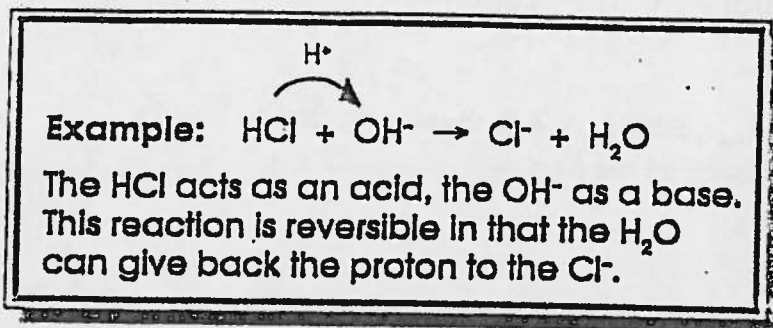


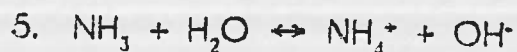
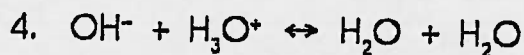
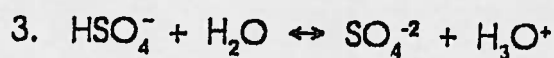
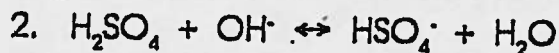
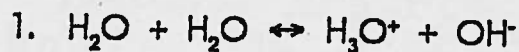
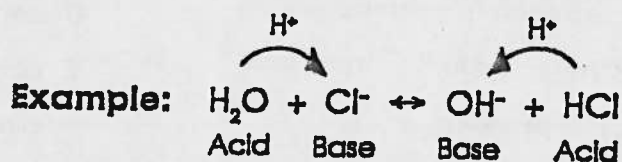
BRONSTED-LOWRY ACIDS AND BASES

Name _____

According to Bronsted-Lowry theory, an acid is a proton (H^+) donor, and a base is a proton acceptor.



Label the Bronsted-Lowry acids and bases in the following reactions and show the direction of proton transfer.





Name _____

Date _____

Period _____

Workbook Activity, Page 2

Chapter 15

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Chapter 15 Vocabulary Review, continued

Directions Match the items in column A with those in column B. Write the letter of each correct answer on the line.

Column A

- _____ 11. proton donor
- _____ 12. molecule or ion formed when an acid donates a proton
- _____ 13. has a known concentration in a titration
- _____ 14. produced from a strong acid and a strong base
- _____ 15. completely dissociates in water
- _____ 16. produced from a weak acid and a strong base
- _____ 17. point when all of the acid reacts with all of the base
- _____ 18. does not completely dissociate in water
- _____ 19. proton acceptor
- _____ 20. uses a scale of 0 to 14
- _____ 21. molecule or ion formed when a base accepts a proton

Column B

- A acid
- B base
- C basic salt
- D conjugate acid
- E conjugate base
- F equivalence point
- G neutral salt
- H pH
- I standard solution
- J strong base
- K weak acid

